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<i>Physical Sciences</i>		<i>CAPS</i>	<i>2014, June Examinations</i>
QUESTION 5			
5.1	5.1.1 Aufbau diagram ✓✓		(2)
	5.1.2 Sulphur ✓✓		(2)
	5.1.3 $1s^2 2s^2 2p^6 3s^2 3p^2$ ✓✓		(2)
	5.1.4 Group 16 ✓ Period 3 ✓		(2)
	5.1.5 The element has 6 valence electrons ✓ and belong to the third horizontal row of the periodic table ✓		(2)
	5.1.5 Non-metal ✓		(1)
5.2	5.2.1 5 and 1, respectively		(1)
	5.2.2 $\begin{array}{c} \text{Y} \\ \text{X} \\ \text{Y} \end{array}$ ✓ "✓ Y		(2)
	5.2.4 N ✓ and H ✓, or P and H		(2)
			(2)
			[18]
QUESTION 6			
6.1	(a) 24 ✓ (b) 12 ✓ (c) Na ✓ (d) 10 ✓ (e) 12 ✓ (f) 17 ✓ (g) 19 ✓		(7)
6.2	Relative atomic mass = $(0,78 \times 35)$ ✓ + $(0,24 \times 37)$ ✓ = 35,48 ✓		(3)
6.3	Isotopes ✓✓		(2)
			[12]
QUESTION 7			
7.1	Sodium hydroxide (NaOH) ✓ Hydrogen ✓ or $H_{2(g)}$		(2)
7.2	$2Mg(s) + O_2(g) \rightarrow 2MgO(s)$ ✓✓ (bal)		(3)
7.3	Sodium + water → sodium hydroxide + water $2Na(s) + 2H_2O(l) \rightarrow 2NaOH(aq) + H_2(g)$ ✓✓ (Bal)		(5)
7.4	Because they react vigorously with air to form oxides ✓ and with water to form hydroxides ✓		(2)
7.5	It is an inert colourless gas ✓ It is unreactive ✓		(2)
7.6	Noble gases		(1)
7.7	Because they don't react, it was not easy to discover them, so no one knew they existed ✓✓		(2)
<i>Limpopo Province</i>		<i>Grade 10 Paper 2 MEMO</i>	<i>Page 4 of 5</i>

Table of Contents

Grade 11 November Agricultural Science 2016 Memorandum	3
Agricultural Science Paper 2 Grade 11	4
Grade 11 Exam P1 Agricultural Sciences November 2016	5
Grade 11 November Agricultural Science P1 Memorandum	6
Memo Agricultural Sciences November 2013	7
Agricultural Science P1 Memorandum	8
Paper2 Life Science Grade11 Caps Memorandum	9
Memorandum Agricultural Science Paper	10
Agricultural Science Paper On Memorandum 2014 Grade 10	11
Agricultural Sciences Grade 12 Past Papers And Memos From	12
Agriculture Paper 2 Grade12 2013 November Memo	13
Agricultural Science Paper1 Memo 2013 November	14
Grade11 Agricultural Science Questions Paper1 And	15
Agric P1 Nsc Dbe November 2014 Memo	16
Agricultural Science Memo For 2018 November Grade 12	17
Agricultural Sciences Paper1 Memorandum Limpopo September 2012	18
Eastern Cape Grade 11 Past Exam Papers	19
Memorandum Of Agriculatural Science Paper2 November 2014	20
Paper 2 Agricultural Science Nsc 2013	21
Limpopo Grade 11 Agriculture Memorandum Paper1	22
.....	23

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